

Multiple Choice Question For Molarity Of Solution

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MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question. 1) Formation of solutions where the process is endothermic can be spontaneous provided that _____. A) the solvent is a gas and the solute is a solid B) they are accompanied by an increase in order C) they are accompanied by an increase in disorder

AP Chem: Chapter 4 Practice Multiple Choice Questions Multiple Choice Identify the choice that best completes the statement or answers the question. Calculate the molarity of the resulting solution if 25.0 mL of 2.40 M HCl solution is diluted to 300. mL. a. 0.200 M. b. 29.0 M

come out of solution. 8. Molarity is a temperature dependent quantity, whereas molality is not. Explain why this is so. Molarity depends upon the volume of the solution and this property changes (generally rather slightly with liquid solvents) with temperature. Molality depends on mass which is not temperature dependent. 9.

MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question. 3) Of the species below, only _____ is not an electrolyte. A) HCl B) Ar C) NaCl D) KOH E) Rb₂SO₄ 4) What is the molarity of an aqueous solution containing 75.3 g of glucose (C₆H₁₂O₆) in 35.5 mL of solution? A) 0.197 B) 2.12 C) 3.52 D) 1.85 E) 11.8

MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question. 1) The molarity of a solution prepared by diluting 43.72 mL of 5.005 M aqueous K₂Cr₂O₇ to 500 mL is _____. A) 0.0879 B) 0.0044 C) 57.2 D) 0.870 E) 0.438 1) Answer: E 2) Of the species below, only _____ is NOT an electrolyte.

Find the molarity of a solution containing 60g of BaCl₂ in 320 mL of solution. (3 points) 6. How many milliliters of an 18 M solution of acetic acid (HC₂H₃O₂) would you need to make 575 mL of a 2.0 M solution **MULTIPLE CHOICE:** Choose the letter that best answers the statement or question. (2 points each) 1. A solution that contains as much

CHEMISTRY Three hours are allotted for this examination. One hour and 30 minutes are allotted for Section I, which consists of multiple-choice questions. For Section II, Part A, 40 minutes are allotted; for Section II, Part B, 50 minutes are allotted. The molarity of the solution remains unchanged. (C) The molality of the solution remains

Part A – Multiple choice test Choose one correct answer for each question and write corresponding letter in answer sheet provided. Question 1 solution is a) 0.1 M, b) 1.0 M and 5.0 M.

Sample Exam 5 (Solutions) Multiple Choice 1. Which colligative property measurement is best to use for compounds with a molar mass greater than 5000 g/mol? a. osmotic pressure A crystal is placed in a solution and it dissolves. The solution must have been a. unsaturated.

d. They increase the concentration of hydroxide ions in aqueous solution. ____ 14. Strong bases are a. strong electrolytes. c. nonelectrolytes. b. weak electrolytes. d. also strong acids. ____ 15. Whose definition of acids and bases emphasizes the role of protons? a. Brønsted and Lowry c. Arrhenius b. Lewis d. Faraday ____ 16. An electron-pair